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problem gives us,  
for a hypothetical  
ion pair, values for  
 $r_0$  (0.35 nm),  $E_0$   
6.13 eV), and  $n$

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(10), and asks that we determine explicit expressions for attractive and repulsive energies of Equations 2.8 and 2.9. In essence, it is necessary to compute the values of A and B in these equations.

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for this metal at  
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This problem is  
solved using two  
steps: (1) calculate  
the Callister

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700°C for 20 h the specimen in part (a) will transform the coarse pearlite and martensite to spheroidite. (c) Rapidly cool to 600°C (1110°F), hold for 4 s, rapidly cool to

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Solution This  
problem is solved  
using two steps: (1)  
calculate the total  
number of lattice

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Sites in silver,  $N_{Ag}$ , using Equation 4.2; and (2) multiply this number by fraction of lattice that are vacant,  $2 \times 10^{-6}$ . The parameter  $N_{Ag}$  is related to the density,  $(\rho)$ , Avogadro's number ( $N_A$ ), and the atomic weight ( $A_{Ag} = 107.87 \text{ g/mol}$ ,

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Using two steps: (1) calculate the total number of lattice sites in silver,  $N_{\text{Ag}}$ , using Equation 4.2; and (2) multiply this number by fraction of lattice that are vacant, 2

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The average atomic weight of silicon ( $A$ ) is computed by adding fraction-of-occurrence/atomic weight Si products for the three isotopes—i.e., using Equation 2.2. (Remember: fraction of occurrence is equal to the percent of

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occurrence divided  
by 100.) Thus. A Si  
= f 28 A 28 + f 29  
A 29 f 30 A Si Si Si  
Si Si 30 Si

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Binary Nickel Alloys  
, P. Nash (Ed.), ASM  
International,  
Materials Park, OH  
(1991). □ 2 phases:

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L (liquid)  $\alpha$  (FCC  
solid solution) □ 3  
phase fields: L L +  
 $\alpha$   $\alpha$  0 20 40 60 80  
100 wt% Ni 1000  
1100 1200 1300  
1400 1500 1600  
T(°C) L (liquid)  $\alpha$   
(FCC solid ...

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researchers to think about since we are used to doing this daily. For our colleagues, it might not come as naturally. Luckily, there are a few different methods we can employ to help us reframe the solution.

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