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atoms are there in 2.00 g of boron?

Solution ANALYZE What is given in the problem? mass of boron in grams

What are you asked to find? number of boron atoms PLAN What steps are needed to convert from grams of B to

...

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$g(x) = 6(0.2x) + 5$ . Select the correct choice below, and fill in the answer box to complete your choice. A. The graph of  $g(x)$  is translated      unit (s) to the left compared to the graph of  $f(x)$ . B. The graph of  $g(x)$  is translated

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unit(s) down compared to graph of  $f(x)$ .

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Is the answer reasonable? Yes; 2 g of  
boron is about  $\frac{1}{5}$  of the molar mass

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of boron. Therefore, 2.00 g boron will contain about  $1/5$  of an Avogadro's constant of atoms. Practice 1.

Calculate the number of atoms in each of the following masses:



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